- 1. The units for the rate constant of a second-order reaction are which of the following
 - a. L mol^-1s^-1
- 2. The units for the rate constant of a zero-order reaction are which of the following
 - a. mol L^-1s^-1
- 3. The rate law of a particular reaction is k[PSO₄]. This means that the reaction is

_____ order and will have a _____ slope.

- a. first, linear
- 4. The rate law of a particular reaction is $k[C_4H_6]^2[H_2O_2]$. This means that the

- reaction is _____ order in C₄H₆ and _____ order overall **CO**. UK a. second, third at any more it in time, determined by the slope 5. The rate of a chemical level oncentration as a function of time defines iential to 🕕

which of the following term

a. instantaneous rate