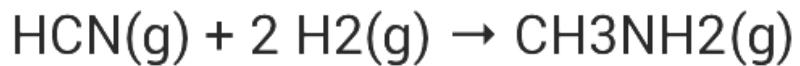


Determine the equilibrium constant for the following reaction at 587 K.



$$\Delta H^\circ = -213 \text{ kJ}; \Delta S^\circ = -188.3 \text{ J/K}$$

- 1.00
- $1.10 \times 10^5$
- $1.31 \times 10^9$
- $6.56 \times 10^9$
- 102.5

Correct

| Preview from Notesale.co.uk  
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