

Determine the equilibrium constant for the following reaction at 587 K.



$$\Delta H^\circ = -213 \text{ kJ}; \Delta S^\circ = -188.3 \text{ J/K}$$

☐ 1.00

☐ 1.10×10^5

☒ 1.31×10^9



Correct

☐ 6.56×10^9

☐ -102.5

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