

★ Atomic Size

It refers to the radius of an atom. It may be visualised as the distance b/w the centre of the nucleus & the outermost shell of an isolated atom.

Variation

decreases across the Period



Increases down the group

Reason for group: New shells are being added as we go down the group. This increases the distance b/w the outermost electrons & the nucleus.

Reason for period: This is due to an increase in nuclear charge which tends to pull the valence electrons closer to the nucleus & reduce the size of the atom.

★ Metallic & Non-metallic character

- Elements having a tendency to lose one more electrons & form positive ions are called metals. (because e^- ion is formed $\therefore$ also called electropositive elements).
- Metals are present in the left side as well as in the centre of the periodic table.
- The tendency of these elements to lose electrons are called metallic character.