If we try to write down the electronic configuration of potassium (Kalium) (K, Z=19) according to above trend, the last electron must go to the 3d subshell, i.e. K19 = [Ne] $3s^2$ $3p^6$ $3d^1$ or [Ar] 3d1 but this electron is said to enter the 4s–subshell according to lower(n+l) rule of Aufbau principle

Sc (Z = 21) [Ar] 4s2 3d1

Ti (Z = 22) [Ar] 4s2 3d2

V (Z = 23) [Ar] 4s2 3d3

Cr (Z = 24) [Ar] 4s1 3d5

Mn (Z = 25) [Ar] 4s2 3d5

Fe (Z = 26) [Ar] 4s2 3d6

Co (Z = 27) [Ar] 4s2 3d7

Ni (Z = 28) [Ar] 4s2 3d8

Cu (Z = 29) [Ar] 4s2 3d10

Zn (Z = 30) [Ar] 4s23d10

The next six elements, *that is*, Ga31 to Kr36 belong to p-block and the last electron enters the 4p – subshell of the atoms of these elements. The electronic configurations of these elements are as follows:

Gallium Ga (Z = 31) [Ar] 4s2 3d10 4p1

Germanium Ge (Z = 32) [Ar] 4s2 3d10 4p2

Arsenic As (Z = 33) [Ar] 4s2 3d10 4p3

Selenium Se (Z = 34) [Ar] 4s2 3d10 4p4

Bromine Br (Z = 35) [Ar] 4s2 3d10 4p5

Krypton Kr (Z = 36) [Ar] 4s2 $3d^{10}$ 4p6 or 1s2 2s2 2p6 3s2 $3p^6$ $3d^{10}$ 4s2 $4p^6$

In next two elements of 5th period, the last electron goes to 5s- subshell. (s-block):

Rubedium Rb (Z = 37) [Kr] 5s1

Strontium Sr (Z = 38) [Kr] $5s^2$