

hinders the direct attraction between the nucleus and electrons. So, electrons are held loosely. The electronic configuration is shown below.

Element	Atomic Number	Electronic Configuration
Beryllium (Be)	4	[He]2s ²
Magnesium (Mg)	12	[Ne]3s ²
Calcium (Ca)	20	[Ar]4s ²
Strontium (Sr)	38	[Kr]5s ²
Barium (Ba)	56	[Xe]6s ²
Radium (Ra)	88	[Rn]7s ²

As it is clear from the table. All alkaline earth metals contain two valence electrons. So, they have similar chemical and physical properties.

Properties of Alkaline Earth Metals

Properties of alkali metals are discussed below.

Size

Alkali earth metals have a smaller radius than group 1 in the periodic table. As moving down the group, the number of shells increases. Due to this atomic radius increases as moved down the group.