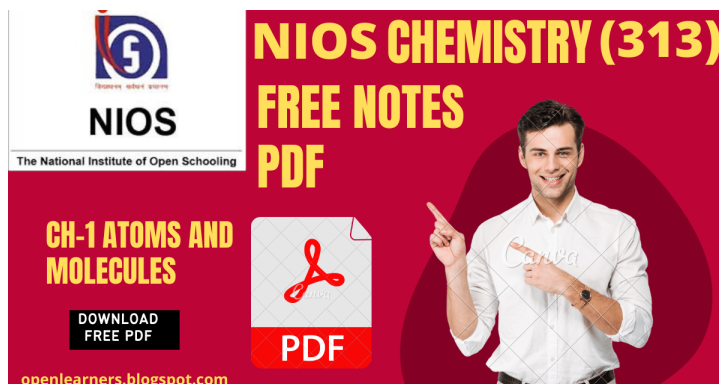


Module 1 Some basic concepts of chemistry.
Ch -1 Atoms , molecules and chemical arithmetic.



Note :- this pdf is specially for quick revision it contains important topics which have more chances that they are coming in final exams.

Law of conservation of mass.

In every chemical reactions, total masses of all the reactants is equal to the masses of all the products.

Law of definite proportion.

In a given chemical compound the percentage of mass remains constant either it is natural or synthetic.

Law of multiple proportion.

Those reactions in which reactants are same but products are different then if the mass of the one reactant is constant then mass of other reactants in the simple ratio.

Avogadro's law

Equal volume of gases under STP (standard temperature pressure) contain equal no. of molecule and atoms.

Dalton's Atomic Theory (very important)

Postulates of Dalton's Atomic Theory

Dalton's Atomic Theory's four postulates is related to the law of conservation of mass.

1. Matter consists of indivisible atoms.
2. All the atoms of a given chemical elements are identical in mass and in all other properties.
3. Atoms are indestructible and remain their identify in chemical reactions.