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If two elements are produced at an electrode, the less reactive element will usually be formed.

The electrolysis of brine (sodium chloride solution) is an important industrial process.

Aqueous solution - When a substance is dissolved in water.

Bases - Chemicals which can neutralise acids.

Alkalies - Bases which can dissolve in water.

- All acids taste very sour. All acids form H^+ ions when added to water - the hydrogen ions make a solution acidic.

- Bases are the opposite of acids in the way they react. All alkalies form hydroxide ions (OH^-) when added to the water, this makes the solution alkaline.

Indicators - special chemicals which change colour when added to acids and alkalis.

pH Scale - to show how acid/alkaline a solution is. (0 most acidic, 14 most alkaline).

Common indicators: Litmus paper, Phenolphthalein, Methyl Orange and universal indicator.

$\text{Metal} + \text{Acid} \rightarrow \text{Salt} + \text{Hydrogen}$.

$\text{Base} + \text{Acid} \rightarrow \text{Salt} + \text{Water}$.

$\text{Carbonate} + \text{Acid} \rightarrow \text{Salt} + CO_2 + \text{Water}$.

Precipitates Precipitation reactions are used to remove pollutants from the wastewater from factories.