

DEFINITIONS

THERMODYNAMICS

is the scientific study of the interconversion of heat and other kinds of energy.

THERMOCHEMISTRY

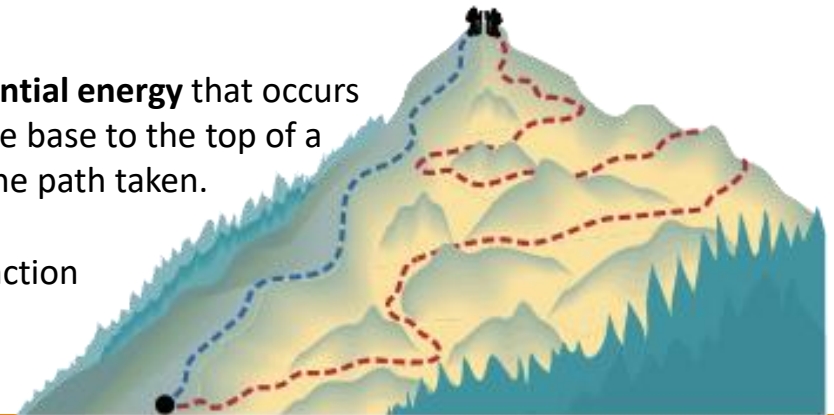
is the study of *heat change* in chemical reactions.

FUNCTIONS OF A SYSTEM

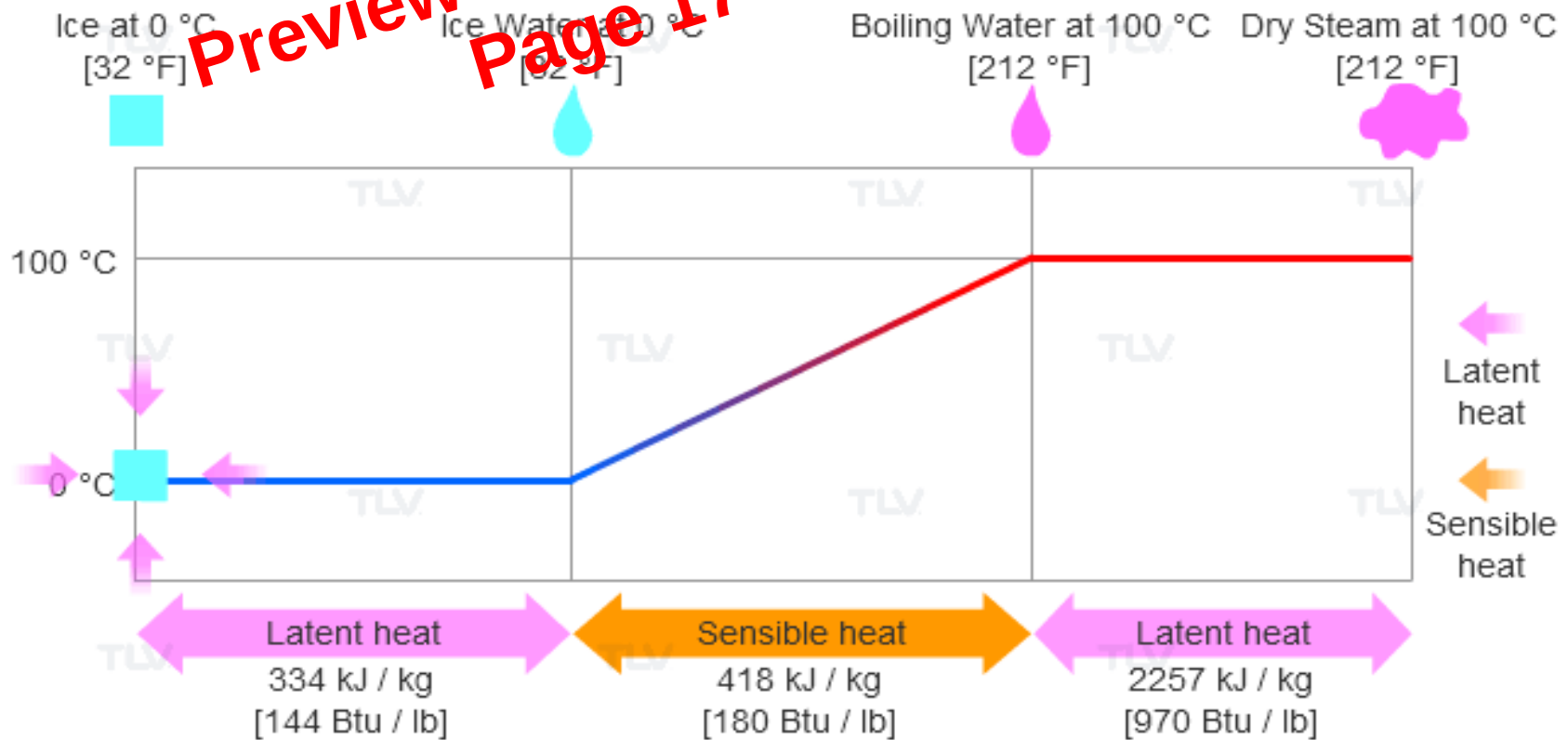
- **State functions** are properties that are determined by the state of the system regardless of how that condition was achieved. These depend only on **the initial and final states of the system** and not on how the change is accomplished. (i.e. energy, pressure, volume, temperature)
- **Path functions** are properties or quantities whose values depend on the **transition of a system** from the initial state to the final state (pathway). (i.e. work, heat)

The **gain in gravitational potential energy** that occurs when a person climbs from the base to the top of a mountain is independent of the path taken.

Potential energy is a state function
Work is a path function



LATENT HEAT AND SENSIBLE HEAT



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