magnetic field. This is the outcome of whom's rule of maximum multiplicity. The electronic configuration is governed by the three concept of bow principle which says that electrons are filled in the increasing order of energy according to n plus I rule multiplicity. It says that degenerate orbitals will get field first singly and then pairing will take place and according to extra stability principle half or completely fields options are extra stable and so that configuration is preferred so d5 configuration will be more stable as compared to d4 configuration.

We will be discussing a few questions that may arise in the minds of the students and also they are there in the examination. The Balmer series is obtained when an electron jumps from higher energy levels to second energy level. Ozone molecules have lambda values that are inversely proportional to their molecular masses. If they are moving with the same velocities, then we can calculate the wavelength of the Balmer line. Z square upon n square is the nuclear charge or atomic number. We can say n is the number of the shell and the negative sign indicates that the energy is less by this publicular factor as compared to the energy of the free electron. When the electron is bound to the nucleus in an atom, there will definitely be negative sign for hydrogen atom.

For n equal to 21 Onhave value 0 by Or lequal to 0, m can not have value equal to minus 2. To mis is not a correct quantum number in this particular group. m can have values from plus I2 minus I to minus I that is plus 1 to plus 1 so minus 1 value is a valid value. There are three invalid sets of quantum numbers and there are two valid quantum numbers. Now let us see what is the angular momentum of an electron revolving in the p orbital. If an electron is revolving in a p orbital which is having a dumbbell shape, what is angular momentum? Let us see. Let us calculate the number of nodal planes in various orbitals. There is no minima in between two maximas coming to the 5 p y orbital for phi u p by orbital we know that for 5 p subshell n equal to 5 and I equal to 1 and hence n minus I minus 1 that is equal to number of planes comes out to be equal to phi u minus 1 minus 1 equal to 3.

The electron magnetism is because electron is continuously revolving around its own axis it is behaving as a bar magnet. When an electron is placed in an orbital having I equal to 2 then it can align in five different ways in one case there is no effect in second case there will be plus half plus h cross plus twice h cross minus h cross and so on. Let us write first the electronic configuration of iron with atomic number. Li positive ion has