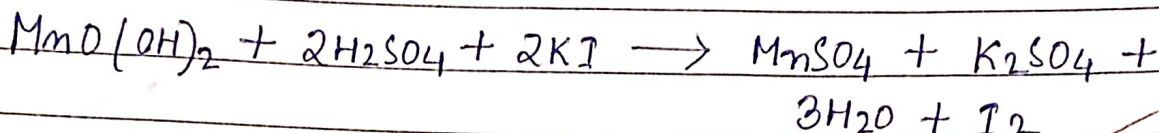
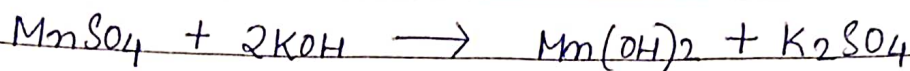


Aim of the Experiment :- To determine dissolved oxygen in water sample (tap water)

Theory :-



Liberated iodine in an amount equivalent to the original dissolved oxygen content of the sample. The I_2 is determined by (liberated) titrated with $\text{Na}_2\text{S}_2\text{O}_3$.

Material and Reagent :- Conical flask of 250 ml, Pipette, MnSO_4 , Potassium iodide, potassium hydroxide, Sodium thiosulphate, starch indicator, HCl, Conc. H_2SO_4

Procedure :-

(A) Standardisation of $\text{Na}_2\text{S}_2\text{O}_3$:- (i) Fill the burette with a solution of $\text{Na}_2\text{S}_2\text{O}_3$.

(ii) In a conical flask of 250 ml add 50 ml of 0.01N of $\text{K}_2\text{Cr}_2\text{O}_7$ (0.1g in 250 ml water).

(iii) Add 1.5 ml of 2M HCl followed by 1g potassium iodide.

(iv) Now, the iodine liberated is titrated against $\text{Na}_2\text{S}_2\text{O}_3$ till the solution becomes pale yellow.