- Test for HCl gas with litmus paper, **DON'T** give the same result as Cl₂ gas.
 At this time, you should know the results of these tests for HCl and Cl₂ gases (please see pg 20 of Chapter 1 Acid, Base & Salt module)
- Difficulty in explaining the definition of Rate Of Reaction, ROR:
 It's A change in selected quantity (change in mass/volume) over a fixed period of time.
 Students always omit the sentence in bold!
- 12. Remember that the total volume of gas released by the same amount (mass/size/length) of metal is always the same.
 <u>A common mistake</u> is to say that the powdered metal, when reacted with acid, gives off more gas than larger lumps of the same amount of metal.
 Volume of gas released/ given off should be the same. The difference is the powdered metal, when reacted with acid has a higher ROR!!
 13. The total
- 13. The total volume of gas released by a **citized** reaction is exactly the **same** as for an uncatalysed reaction. The same amount of peaceans is the important factor!
- 14. For ROR questions, when asked to analyse graphs of volume of gas against time for the reaction of acid with a metal/metal carbonate, a common error is mentioning the volume of gas is increasing and not mentioning the ROR. Remember that the ROR is decreasing with time as the reaction proceeds because ROR = $\frac{difference in volume of gas}{time taken}$
- 15. What does the symbol "aq" mean?

Answer "aqueous" alone isn't enough to fetch you 2 marks! You must include Solution ✓ formed when solute dissolved in water ✓

Additional info: bromine at room temperature is liquid but once it's displaced from KBr (aq) by Cl₂ (g) {see pg 11 of Yr10 Chapter 4 Periodic Table module} then bromine will be in aqueous state. If bromine liquid is diluted with water, then bromine will be in aqueous state.