

C3 L5 : The Importance Of Alkalis

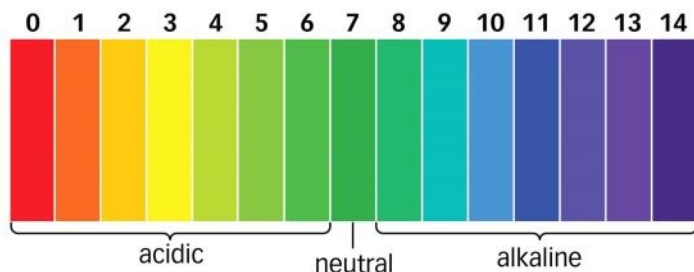
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Traditional sources and its uses

Recall characteristics of alkaline hydroxides and carbonates

large scale methods

neutralizes acids to make salts



Red - 1 - acid

Green - 7 - neutral

Purple - 10 - green

Acids release hydrogen ions (+) into solution - pH (Hence the hydrogen)

Alkalis release OH⁻ into solution (Hydroxide ion)

H⁺ from acids reacts with OH⁻ from alkalis to form H₂O

uses : (For alkalis)

- Adjusting soil pH
- antacid tablets
- soap and cleaners
- making glass
- dying fabric
- paper production

alkali	dissolve in water?	pH of the water	reaction with acid	pH of acid after reacting with alkali	reaction with limewater
Metal hydroxide (Calcium hydroxide)	no	10	no	1 -> 10	
Metal carbonate (Sodium carbonate)	yes	10	yes	1-> 7 neutralise	cloudy - releases CO ₂