

Electrochemistry

Faraday's Ist law

$$m \propto d, \quad m = ZQ$$

$$I = \frac{Q}{t}, \quad Q = It$$

Z: Electro Chemical equivalent

$$1F = 96,500 C$$

$$Z = \frac{\text{eq. mass of a substance}}{96500}$$

$$\text{eq. mass} = \frac{\text{molar mass}}{n\text{-factor}}$$

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$$m = ZIt$$

$$m = \frac{\text{molar mass} \times It}{n\text{-factor} \times 96500}$$

Faraday's IInd law

$$\frac{\text{mass of 1}}{\text{Eq. mass of 1}} = \frac{\text{mass of 2}}{\text{Eq. mass of 2}} = \dots$$

Conductivity of cell

Conductance (G)

$$G = \frac{1}{R} \text{ (ohm}^{-1} \text{ or S)}$$