

- Potassium
- Sodium
- Lithium
- Calcium
- Magnesium
- Aluminium
- Iron
- Zinc
- Copper
- Silver
- Gold

Distillation
 cement
 energy
 melting
 sublimation

Particle	Location	Relative mass	Relative charge
Proton	Nucleus	1	+1
Neutron	Nucleus	1	0
Electron	shells	$\frac{1}{1836}$	-2

atomic mass number (protons + neutrons)
 $23 \leftarrow$
 $11 \leftarrow$ Na
 atomic number (protons)
 Mendeleev $\rightarrow$ 289

metals, transition metals, non metals
 arranged in order of increasing
 atomic number

Mass = moles $\times$ Mr
 Moles = $\frac{\text{mass}}{\text{Mr}}$

Chemistry

Reactivity series
 Acids + Bases
 Titrations
 Separation techniques

Universal indicator

Indicator	Acid	Neutral	Base
Universal	Red	Yellow	Blue
Phenolphthalein	Colourless	Colourless	Pink

Indicators

paper chromatography
 separating points

simple distillation
 CO₂ 0.04%
 Oxygen 21%
 Nitrogen 78%
 Argon 0.9%
 (AR)

fractional distillation
 BP is too similar

Bonding
 ionic
 Covalent $\rightarrow$ Share
 Simple covalent
 giant covalent
 delocalised
 metallic
 ionic structures
 covalent structures

Chemistry

Atomic structure
 Periodic table
 Chemical calculations

Isotopes
 Halogens
 Alkali metals
 Transition metals
 Group 1
 Group 2
 Group 7
 Group 0

metals, transition metals, non metals
 arranged in order of increasing
 atomic number

Mass = moles $\times$ Mr
 Moles = $\frac{\text{mass}}{\text{Mr}}$