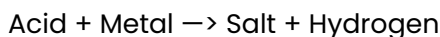


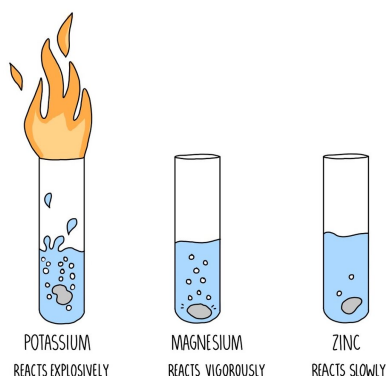
Reactions of Metals



You can use the reaction of different metal with dilute acids to work out how reactive they are

The most reactive metals will react explosively e.g. sodium. The speed of the reaction is indicated by the rate of bubbles produced (the amount of hydrogen given off).

Hydrogen can be confirmed by the **burning splint test**



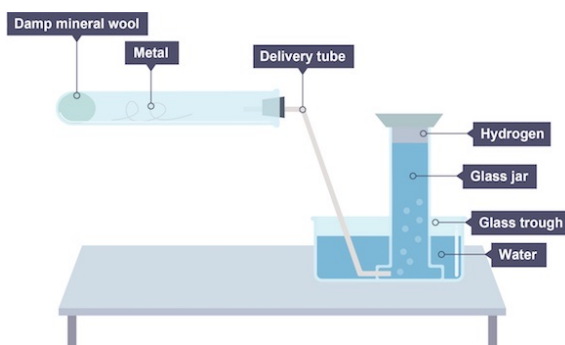
The potassium should give off the loudest 'squeaky pop' as it has the most vigorous reaction producing the most hydrogen gas.



Very reactive metals like potassium, sodium, lithium and calcium all react vigorously with water. The amount of energy given out by the reaction increases with the reactivity of the metal.

Very reactive metals such as potassium, have the ability to ignite hydrogen.

Less reactive metals like magnesium, zinc and iron will not react with water but will with steam.



The mineral wool is heated to create steam.

This experiment cannot be done with aluminium because it does not react with water. It forms a protective layer of aluminium oxide which prevents water from reaching the metal below.

Copper does not react with steam or water