

Orbital hybridisation

- Combining different types of atomic orbitals (s, p, d) in order to create hybridised orbitals
- Hybridised orbitals have different energy/shape from original ones

Rules for hybridisation

- Amount of orbitals does not change during hybridisation
- Hybrid orbitals are all equal

Example:

Hybridisation	Atomic orbitals hybridised	Result
SP	s, p	2 HO sp
SP^2	s, 2* p	3 HO sp^2
SP^3	s, 3* p	4 HO sp^3

