## **Orbital hybridisation**

- Combining different types of atomic orbitals (s, p, d) in order to create hybridised orbitals
- Hybridised orbitals have different energy/shape from original ones

Rules for hybridisation

- 1. Amount of orbitals does not change during hybridisation
- 2. Hybrid orbitals are all equal

## Example:

Hybridisation	Atomic orbitals hybridised	Result
SP	s, p	2 HO sp
SP <sup>2</sup>	s, 2* p	3 HO sp <sup>2</sup>
SP <sup>3</sup>	s, 3* p	4 HO sp <sup>3</sup>

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