

Atoms/Molecules and Stoichiometry

Atomic mass - weighted average mass of an atom

Molecular mass - mass of an molecule

Formula mass - mass of one formula unit of a compound.

Isotope mass - mass of a particular isotope of an element compared with ^{12}C .

The mole amount of substance that has the same number of particles.

Avogadro's constant: 6.02×10^{23}

Mass spectra

Abundance of isotopes can be represented on a mass spectra diagram.

$$\text{relative abundance} = \frac{\text{Peak height}}{\text{total height}} \times 100$$

$$A_r = \sum \frac{\text{mass} \times \text{relative abundance}}{100}$$