Q.2 What type graph will you get when PV is plotted against P at constant temperature.

Q.3 What would have happened to the gas if the molecular collisions were not elastic?

Q.4 At a particular temperature, why vapour pressure of acetone is less than of ether?

Q.5 Why liquids diffuse slowly as compared to gases?

Q.6 What would be the SI unit for quantity $\frac{P.V^2.T^2}{n}$?

Q.7 In terms of Charle's law explain why -273^o C is the lowest temperature?

Q.8 For real gases the relation b/w P,V,T is given by vander Waal's equation, write it for n moles?

Q.9 What correction is applied to obtain pressure of dry gas with the total pressure & aqueous tension?

Q.10 Name two phenomena that can be explained on the basis of surface tension.

Answers to two marks questions

Ans 1. 1atm = 101325 Pa or Nm^{-2}, 1 bar = 10^5 Pa.

Ans 2 .A straight line parallel to pressure axis.

Ans 3. On every collision there is loss of energy, so molecules would have slowed down & settled down in vessel and pressure reduce to O.

Ans 4. b/c molecular force of attraction in acetone is stronger than those present in ether.