The absorption spectra of apple juice reveals that

- A.) Apple doesn't contribute to the deep red color of cranberry-apple juice
- B.) Apple is responsible to the deep red color of cranberry-apple juice
- C.) None of the above
- D.) Not possible to predict ANSWER- A

What is true about the calibration curve for this experiment A.) Should include all tean trendline PIEP Page B.) Paraction

B.) Percent cranberry is determined using the equation y=mx+b

C.) Both A and B

D.) None of the above - ANSWER- C

In this experiment, the apparatus needed is

A.) Volumetric flasks

- B.) Atomic Absorption Instrument
- C.) GC-Mass instrument
- D.) None of the above ANSWER- C

Which is/are the terms associated with Mass spectrometry

A.) Molecular ion

- C.) Mass to charge ration from Notesale.co.uk Preview Page 13 of 45 D.) All of the above ANSWEE

Which is/are the terms associated with Chromatography

- A.) Mobile phase
- B.) Stationary phase
- C.) Mass to charge ratio

- B.) Atomic Absorption Instrument
- C.) GC-Mass Instrument
- D.) None of the above ANSWER- B

The Atomic Absorption spectroscopy

A.) Measures concentration of vaporized atoms by detecting the absorption of light



With increase in the concentration of Fe, Absorbance

A.) Increases

B.) Decreases

C.) No change

D.) A or C - ANSWER- A

a) barium chloride dihydrate

b) magnesium sulfate heptahydrate ______ - ANSWER-BaCl2. 2H2O

MgSO4 · 7H2O

If 125 grams of magnesium sulfate heptahydrate is completely dehydrated, how many grams of anhydrous magnesium sulfate will remain? - ANSWER- 100% - 51.172% H2O = 48.828% MgSO4

(125 g)(0.48828) = 61.0 g MgSO4



A substance undergoing a chemical reaction. - ANSWER- Reactant

A substance resulting from a chemical reaction. - ANSWER- product

A substance that speeds up a chemical reaction. - ANSWER- catalyst

A relative order of metals arranged in a list according to their ability to undergo a chemical reaction. - ANSWER- Activity series

A solution of a substance dissolved in water. - ANSWER- Aqueous Solution

An insoluble solid substance produced from a reaction in aqueous solution. - ANSWER- precipitate

The solution above a precipitate after insoluble particles separate. - ANSWER-supernatant

AB --> A + B - ANSWER- decomp

A + B --> AB - ANSWER- synthesis

AD (acid) + BC (base) --> AC (salt) + BD (water) - ANSWEFC reatralization AC + B --> BC + A - ANSWEFC Single replacement AC + BD --> AD + BC - ANSWER- Double replacement

List the 4 observations that provide evidence for a chemical reaction. - ANSWER-1. change of temperature

- 2. a change of color
- 3. an emission of gas
- 4.the presence of an odor

The indicator phenolphthalein changes the color of a solution based on pH. Phenolphthalein is

in an acidic solution and [Select] in a basic solution. - ANSWER- clear

pink

Provide the chemical formula for the white smoke produced in reaction A.1 from the data table: - ANSWER- 2 MgO (s)

Provide the chemical formula for the strong odor produced in reaction A.2 from the data table: - ANSWER- 2 SO (g)

Provide the chemical formula for the colorless gas produced in reaction B.1 from the data table: - ANSWER-

Provide the chemical formula for the flame-extinguishing gas produced in reaction

Provide the chemical formula for the grave of produced in reaction C.1 from the data table: - ANSWER-Provide the chemical formula for the grave of produced in reaction C.1 from the data table: - ANSWER-Provide the chemical formula for the colorless gas produced in reaction C.2 from the data table: - ANSWEP

the data table: - ANSWER-

Provide the chemical formula for the white precipitate (ppt) produced in reaction C.3 from the data table: - ANSWER-

Provide the chemical formula for the cream-colored precipitate (ppt) produced in reaction D.1 from the data table: - ANSWER-

Provide the chemical formula for the blue-white precipitate (ppt) produced in reaction D.2 from the data table: - ANSWER-

A sample of hydrogen gas is collected over water at 25 °C and the vapor pressure of water at 25 °C is 24 mm Hg. If the barometer reading is 754 mm Hg, what is the partial pressure of the hydrogen gas? (Refer to Table 1). - ANSWER- 730 mm Hg

A 0.0795 gram sample of magnesium metal reacts with hydrochloric acid to produce 88.5 mL of hydrogen gas at 25 oC and 766 mm Hg. From Table 1, the vapor pressure of water at 25 oC is _____ mm Referring to Example Exercise 1, the newly adjusted pressure is therefore, _____ mm Using the Combined gas Law, we can calculate the corrected (new) volume, which is ______ mL In Liters, this value is ______ L. rom their stoichiometric relationship of Mg and H2, the number of moles of H2 produced is ______ mol. The molar volume of hydrogen is therefore, ______ L/mol. - ANSWER- 24.0 742 79.2 0.0792 3.27 x 10-3 24.2 **Preview from 1000 for 1000 for**

A 0.200-g sample of unknown metal (X) is dropped into hydrochloric acid and releases 80.3 mL of hydrogen gas at STP. Using the Ideal Gas Law, the number of moles of the unknown is ______, yielding a molar mass of ______. Referring to the Periodic Table, the name of the unknown metal is ______ and symbol - ANSWER- .003581

55.85

Iron

Fe

Applying Boyle's Law Concept. As a piston compresses air in a cylinder, the gas pressure increases . - ANSWER- increases