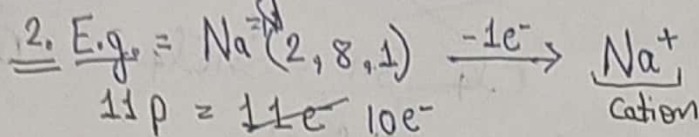


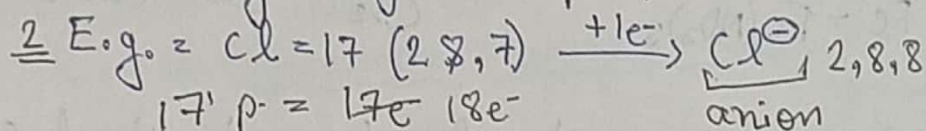
Basics [IOC]

Metals - 1. Generally e^- donors



* Metal Oxide = Basic

Non Metals - 1. Generally e^- gain (acceptor)



* Non Metal Oxide = Acidic (generally)

→ Periodic Table = Metals = LHS, Non metals = RHS
[extreme right - inert gases]

Nomenclature of IOC compound:-

ide - Element alone has a $-ve$ charge
E.g. = H^- , OH^- , O^{2-} , S^{2-} , C^{4-} , X^- etc.
Hydride Hydroxide Oxide Sulphide Fluoride Halide

ic & us forms

[ic > us] E.g. = Fe^{2+} = ferrous, Cu^{+1} = Cuprous, Sn^{2+} = Stannous
 Pb^{2+} = Plumbous, Fe^{3+} = ferric, Cu^{+2} = Cupric, Sn^{4+} = Stannic
 Pb^{4+} = Plumbic

Some imp. anions -

O^{2-} , H^- , C^{4-} , N^{3-} , I^- etc. ide

SO_3^{2-} = Sulphite, SO_4^{2-} = Sulphate, NO_3^- = Nitrate, NO_2^- = Nitrite

CO_3^{2-} = Carbonate, PO_4^{3-} = Phosphate

Ate $\xrightarrow{-O} \text{ite}$ $\xrightarrow{-all O} \text{ide}$

Formula. -

1. Sodium Carbonate = Na_2CO_3

2. Calcium Sulphate = CaSO_4

3. Magnesium carbonate = MgCO_3

4. Lithium Hydroxide = LiOH

5. Lithium nitride = Li_3N

6. Magnesium Nitride = Mg_3N_2

7. Sodium carbonate = Na_2CO_3

8. Potassium nitrate = KNO_3

9. Sodium nitrate = NaNO_3

10. Mag. chloride = MgCl_2

11. Ferric Sulphate = $\text{Fe}_2(\text{SO}_4)_3$