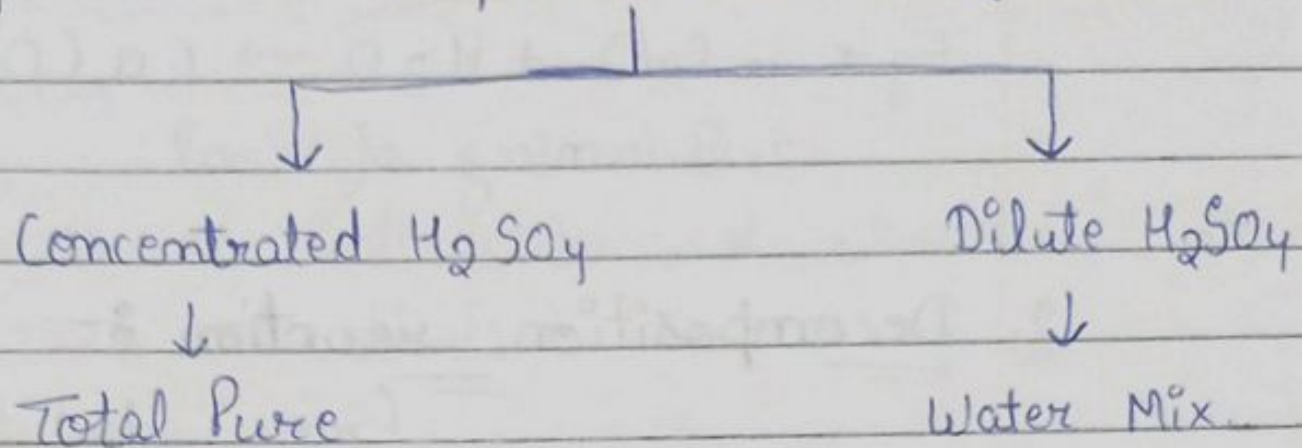
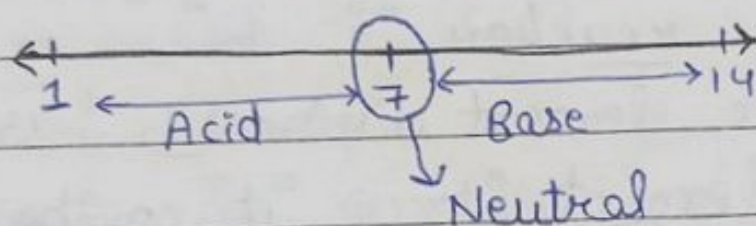


L> Chemical reactions and equations

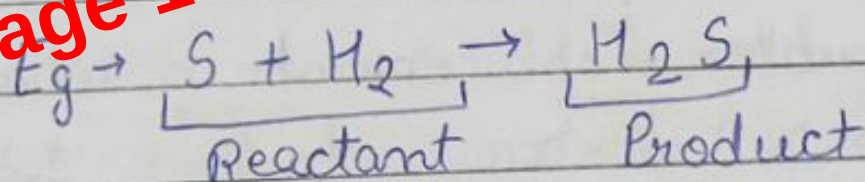
* King of Acids :- Sulphuric acid (H_2SO_4)



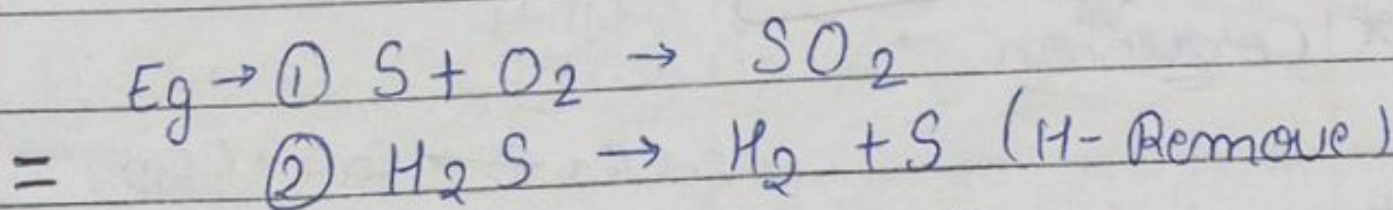
* PH Scale



→ Reduction reaction → किसी reactant के अंदर H-add हो या O-Remove



→ Oxidation → किसी reactant के अंदर H-Remove या O-add



→ Redox → किसी दोनो reactant के अंदर H and O add

Sulphuric acid $\rightarrow$ H_2SO_4
hydrochloric acid $\rightarrow$ HCl .

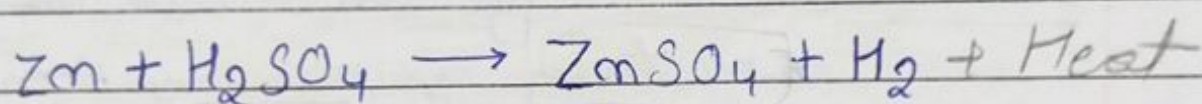
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* ACTIVITY - 1.3 *

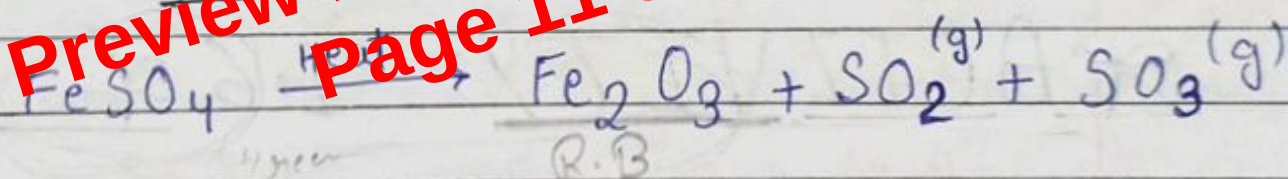
Que-1 What happens when zinc granules are added to HCl/H_2SO_4 ?

Ans- When zinc granules are added to dilute sulphuric acid, there occurs a change in the state of Zn granules and they are converted to Zn sulphate ($ZnSO_4$) and hydrogen gas is evolved.



This reaction is exothermic, i.e. it involves release of heat and so the flask becomes hot and temp. increases.

Que-2. * ACTIVITY - 1.5 *



Que-1 Do you observe any change in the colour of ferrous sulphate crystals when they are heated in the boiling tube over the flame of a burner or spirit lamp?

Ans The green colour of ferrous sulphate crystals changes to brownish black oxide and smell of burning sulphur will be observed.

Ferrous sulphate decomposes to form ferric oxide (Fe_2O_3), sulphur dioxide (SO_2),