

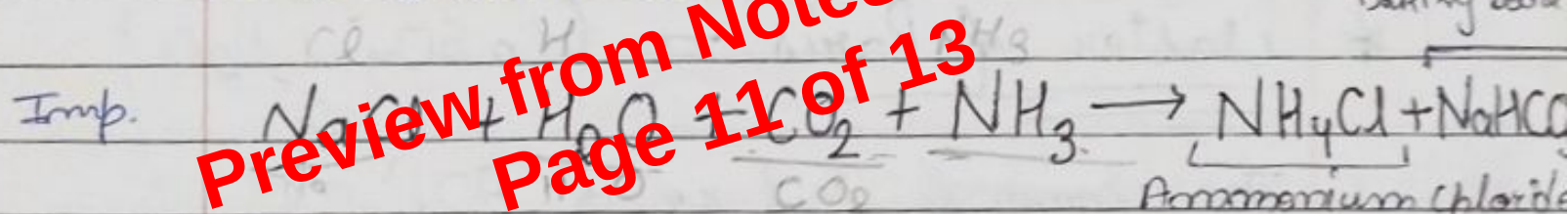
Uses

- Imp. → Used for bleaching cotton
→ used as an oxidizing agent in chemical industries
→ Used for disinfection of water.

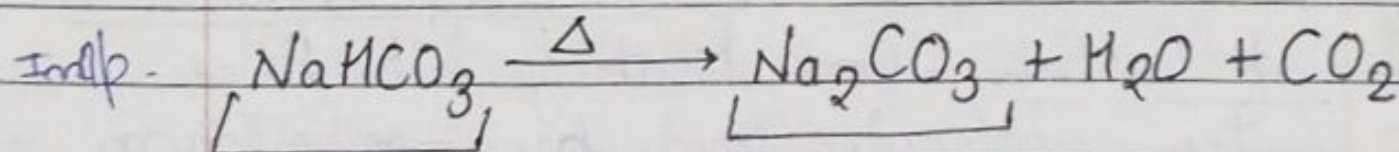
* Baking Soda (NaHCO_3) (Base)

The chemical nature of baking soda is
sodium bicarbonate
Hydrogencarbonate

It is produced on a large scale by treating cold and concentrated solution of sodium chloride with ammonia (NH_3) and carbon dioxide.



On heating, it decomposes to give sodium carbonate with the evolution of carbon dioxide.

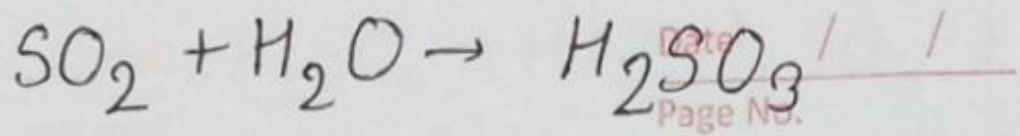


Baking Soda

washing soda

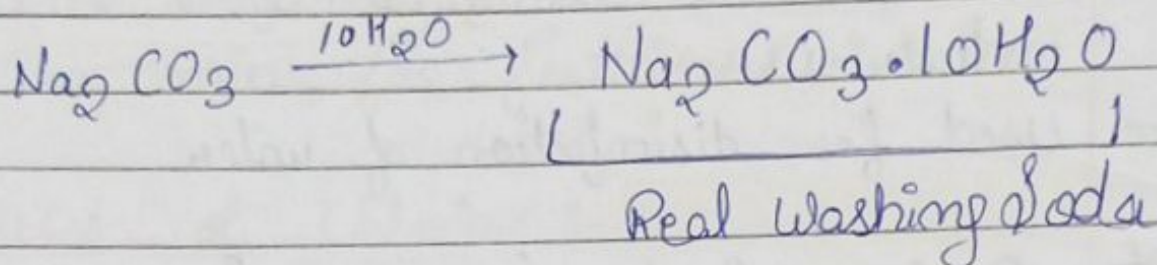
* Uses → (i) Used as an antacid to treat acidity in stomach.

(ii) Used to make baking powder, which is used in preparation of cakes, breads etc.



* Washing Soda :- $(\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O})$

Sodium Carbonate



Uses

1. Used in glass, soap and paper industries
- Imp. 2. Used to remove permanent hardness of water
3. Used in the formation of borax.
 $(\text{Na}_2\text{H}_2\text{OB}_4\text{O}_{17}) \rightarrow$ colorless crystalline solid, that dissolves in water to make a basic solution. (Borax is a salt.)

* Plaster of Paris $(\text{CaSO}_4 \cdot \frac{1}{2}\text{H}_2\text{O})$

The chemical nature of POP -
 Calcium Sulphate Hemihydrate

→ POP is prepared by heating gypsum at 373 K. On heating, it loses water molecules and becomes calcium sulphate hemihydrate.

* When POP absorbs water, it transforms into gypsum

