sp3 and dsp2 are four hybridized orbitals. In any case, one is a tetrahedral shape, and the other is square planar. Why?

sp3 orbitals are framed from the s - subshell with uniform electron dissemination around the core and from the p-subshell with appropriation in the three vertical tomahawks. Hybridized orbitals, subsequently have their electron appropriation in three aspects, as tetrahedral bearings.

In dsp2, all the orbitals associated with hybridization have their electron dispersion around a similar plane. Thus, the hybridized orbitals additionally are in a similar plane leading to square planar math.

The oxygen particle is paramagnetic. Is there a clarification?

An oxygen iota shares two electrons, each with one more oxygen particle, to shape the oxygen atom. Oxygen particles display paramagnetic nature showing unpaired electrons. A sub-atomic orbital hypothesis has been proposed to make sense of this. As indicated by this hypothesis, iotas lose their orbitals and rather structure an equivalent number of orbitals covering the whole atom and thus, the name sub-atomic orbital. Topping off of these orbitals in expanding energy request leaves unpaired electrons making sense of the paramagnetic way of behaving of oxygen atom