

CHEMISTRY

NEET 2023

01 NATURE OF MATTER

08 STOICHIOMETRIC CALCULATIONS

07 LIMITING REACTANT

06 EF & MF

05 PERCENTAGE COMPOSITION

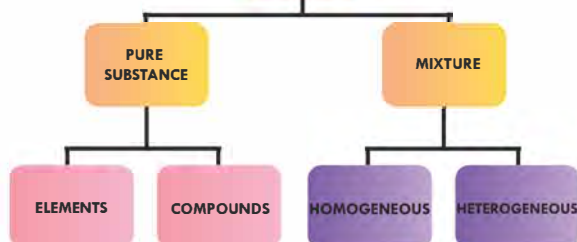
04 MOLE CONCEPT

03 DALTON'S ATOMIC THEORY

02 LAWS OF CHEMICAL COMBINATIONS

SOME BASIC CONCEPTS OF CHEMISTRY

MATTER



01

Which one of the following is not a mixture
(A) Tap water (B) Distilled water
(C) Salt in water (D) Oil in water

LAWS OF CHEMICAL COMBINATION

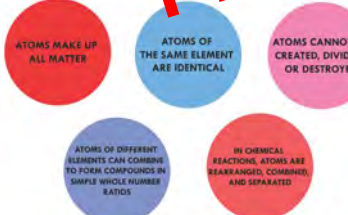


02

Which one of the following pairs of compound illustrate the law of multiple proportions?

- (A) H_2O , Na_2O (B) MgO , Na_2O
(C) Na_2O , BaO (D) $SnCl_2$, $SnCl_4$

03



04

MOLE



Which one of the followings has maximum number of atoms?

- (A) 1 g of $Mg(s)$ [Atomic mass of $Mg = 24$]
(B) 1 g of O_2 [Atomic mass of $O = 16$]
(C) 1 g of $Li(s)$ [Atomic mass of $Li = 7$]
(D) 1 g of $Ag(s)$ [Atomic mass of $Ag = 108$]

NEET 2020

08

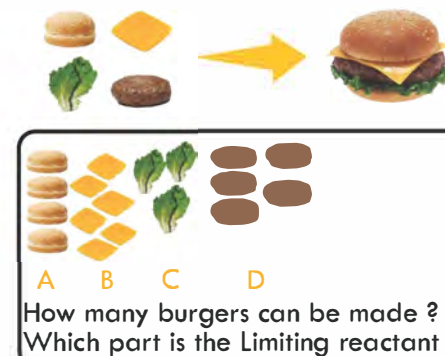
FIRST IDENTIFY THE LIMITING REACTANT BY CALCULATING THE REQUIRED MOLES & AVAILABLE MOLES OF REACTANTS.

THEN CALCULATE THE AMOUNT OF PRODUCT FROM THE AMOUNT OF LIMITING REACTANT

When 22.4L of $H_2(g)$ is mixed with 11.2L of $Cl_2(g)$, each at STP, the moles of $HCl(g)$ formed is equal to
(A) 0.5 (B) 1.5 (C) 1 (D) 2

AIPMT 2014

07



The number of moles of hydrogen molecules required to produce 20 moles of ammonia through Haber's process is

- (A) 40 (B) 10 (C) 20 (D) 30

NEET 2019

06

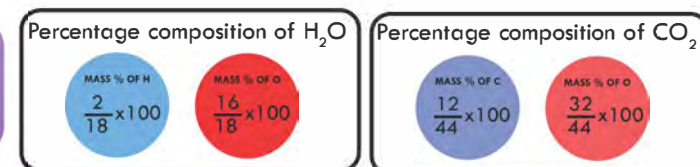
ACTUAL FORMULA Molecular Formula	SIMPLEST FORMULA Empirical Formula
$C_3H_6O_3$	CH_2O
$C_{10}H_{14}N_2$	C_5H_7N
$C_{12}H_{22}O_{11}$	$C_{12}H_{22}O_{11}$

An organic compound contains 80% (by wt.) C & the remaining percentage of H. The empirical formula of this compound is:

- (A) CH_3 (B) CH_4 (C) CH (D) CH_2

NEET 2021

05



Mass % of carbon in ethanol is
(A) 52 (B) 13 (C) 34 (D) 90