

Full Chapter : Basic Concepts of Chemistry

1) Matter

Pure substances

Impure substances

Same type of particles

Different type of Particles

One type of Particles and same Properties

Mixture

They are made up of different types of Particles.

Element Compounds

Made up of one type of Particle

eg. Salt (NaCl), Water (H_2O)

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Mole and Avogadro Number

Relative atomic Mass

Carbon = 12 amu



Here carbon has
only one atom



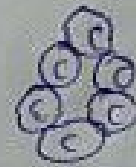
We can't
touch it
see it



So, by solving
this problem
scientists given
the idea of
"Molar Mass"

Molar Mass

Carbon = 12 g/mol



→ total 6.022×10^{23} atoms



Avogadro

Now, we touch it
see it.



One mole = Molar mass of atom
molecule.

So,

Molar of any atom or
molecule = 6.022×10^{23}