

- **Interpretation:** A higher percentage yield means the reaction was more efficient.

Example: If the theoretical yield is 2 moles of water and the actual yield is 1.5 moles:

$$\text{Percentage Yield} = (1.5 \text{ moles} / 2 \text{ moles}) \times 100 = 75\%$$

Summary

- **Theoretical Yield:** Maximum possible product from a reaction.
- **Actual Yield:** Amount of product actually obtained.
- **Percentage Yield:** Efficiency of the reaction; how much product was made compared to the maximum possible.

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Page 2 of 2