

Although, he does not discover proton but later this experiment plays very important role in discovery of proton.

Discovery of electron from cathode rays:

In 1897, British physicist J.J. Thomson conducted experiments using a cathode ray tube. He discovered that cathode rays were deflected by electric and magnetic fields, suggesting that they were composed of negatively charged particles.

Thomson measured the charge-to-mass ratio of these particles and he found that this ratio was much larger than that of any known ion, implying that the particles were much smaller and lighter than atoms and had a high charge.

Based on his findings, Thomson concluded that these particles were a fundamental constituent of atoms. He called them "corpuscles," which were later named "electrons."

Plum pudding theory:

The plum pudding model, proposed by J.J. Thomson, was presented in 1904. He says that atoms were solid structure of uniform positive charge substance with a tiny negative particle stuck inside. It is like plums in pudding.

This model was an early attempt to describe the structure of the atom before more precise models, such as the nuclear model, were developed.



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