

Chem



↳ is an endothermic reaction

- Dilute hydrochloric acid is tested with U.I and with CaCO_3

$\text{pH} \Rightarrow 2$

reaction with $\text{CaCO}_3 \Rightarrow$ colorless gas is given off

→ Cations move to the cathode to gain e^-

↳ and Ions lose e^- at the anode

- * Increased Conc. of CO_2 and methane in air cause an enhanced greenhouse effect

- Jewellery is made of gold alloys rather than pure gold

↳ Alloys are harder

* Zinc carbonate → add aq. NaOH (white ppt)

↳ add dilute HCl (colorless gas)

- the element is less metallic because e^- are lost less easily

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