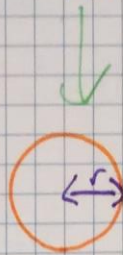
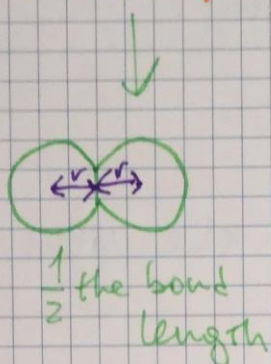
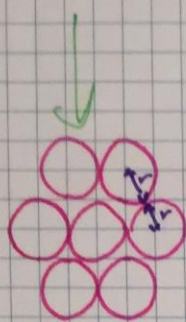


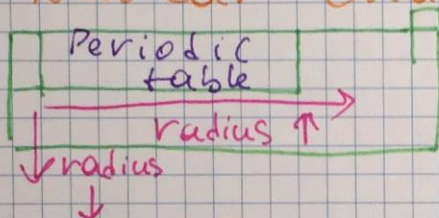
Metallic / Covalent / Vander Waals radius:



Produced by Daniil Flegontov

Factors:

- Shielding effect by the e^- of the inner shells, they repulse outer shells.
- Nuclear charge.



First IF

shielding effect increases down a group

• The further the valence shell, the less are the attractive forces of the nuclear charge.

• Decreases down the group, increases across a period.

Effective nuclear charge

• the nuclear charge experienced by e^- taking into account the shielding.



$$Mg: 12 - 10 = \underline{2+}$$

Decreases with distance.