

Enthalpy Changes in Solutions

- $\Delta H_{\text{sol}}^{\circ}$ is the change in enthalpy when 1 mole of a substance is dissolved in a large excess of a pure solvent
- **$\Delta H_{\text{sol}} = \Delta H_{\text{lattice enthalpy}} + \Delta H_{\text{hydration}}$**
 - Be careful: two ions for $\Delta H_{\text{hydration}}$
- Hydration: when water
- Solvation: anything else

Preview from Notesale.co.uk
Page 6 of 6