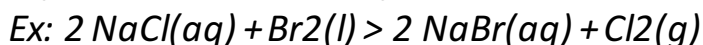


D) +/

Look at the ON of each element in the reactants versus the products. If the ON change, then it is a redox reaction.

Quick check: See if an element when from its elemental form to an ionic form or from an ionic form to its elemental form.



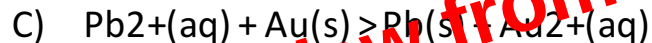
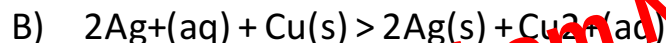
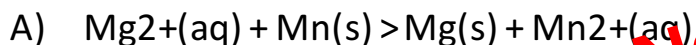
*Br when from 0 to -1. Its ON went down ("**reduced**" / "**oxidizing agent**").*

*Cl when from -1 to 0. Its ON went up ("**oxidized**" / "**reducing agent**").*

Always have to have an element that gets reduced and one that gets oxidized.

Activity Series - used for predicting redox reactions. Elements higher on the list are stronger reducing agent than elements lower down; any metal can reduce the ions of metals below it.

Using the AS, predict which of the following reactions will occur.



Note: If it is higher on the AS, that's the element that would like to go from metallic form to ionic form.

What is the reducing agent in the following redox reaction: $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}$