South Thames College – Centre Number:11058 ID Number:138459

Unit 1

Results:

Mass of Na ₂ CO ₃ and beaker	78.063g
Mass of beaker	76.859g
Mass of Na ₂ CO ₃	1.221g

Analysis:

Moles = Mass/Mr
Concentration=Moles/Volume

Moles=1.221/106
Volume= dm³ = 250cm³ /1000=0.25

Moles= 0.0115mol
Concentration=0.0115/0.25

Concentration of Na2CO3 = 0.046 mol/cm³
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One of the limitations was that no all of the sodium carbonate was transferred from the small beaker in the large beakers (Sme of the sodium carbonate could of remained braces can the remained braces can be always that the balance could be that the bala

remained in the small beaker. Use can be limitation could be that the balance could mislead and alter the measurements because the surface might not be clean and not have any sodium carbonate remaining in it. Also some of the sodium carbonate could of remained at the bottom of the beaker and not all of it was dissolved.

Conclusion:

The calculated concentration of sodium carbonate was found to be 0.046 mol/dm³. The analysis also shows what calculations were used to get the concentration of an unknown solution and in this case it was sodium carbonate solution.