



- Drawing it out as shown indicates that ...
- · the cell reaction goes from left to right
- the electrons go round the external circuit from left to right
- the cell voltage is E°(RHS) E°(LHS). In this way it must be positive
- oxidation takes place at the anode and reduction at the cathode

Conclusion	The reaction(s) will proceed from left to right			
	OX idation RED uction	Zn _(s) > Cu ²⁺ _(aq) + 2e ⁻ >	Zn ²⁺ _(aq) + 2e [−] Cu _(s)	at the ANODE at the CATHODE
	Electrons	Go from the anode to the cathode via the external circuit		
	Cell reaction	Zn _(s) + Cu ²⁺ _(aq) >	Zn ²⁺ (aq) + Cu(s)	