

atom shows if it has been **oxidized** or **reduced** during a reaction.

Oxidation; "an increase in oxidation number" or losing electrons and becoming more positive.

Reduction; "a decrease in oxidation number" or gaining electrons and becoming more negative.

Example

- $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$
- The Na starts with an oxidation number of 0 (It is in its elemental form and is not an ion). Once bonded in the compound NaCl, its oxidation number is +1. It has **increased** in oxidation number and so has been **oxidized**.
- The Cl starts with an oxidation number of 0 (It is also in its elemental form and is not an ion). Once bonded, its oxidation number is -1. It has **decreased** in oxidation number and so has been **reduced**.