

FORM 3 ANALYTICAL EQUATIONS

NAME _____

Complete the equations below. and describe the missing tests and results.

SALT	CATIONS		ANIONS	
	NaOH	Ammonia Solution	Other test	Result
Calcium Nitrate	$\text{Ca}^{2+}(\text{aq}) + \text{OH}^{-} \rightarrow \text{Ca}(\text{OH})_2(\text{s})$	Same as with NaOH	add aqueous ammonia aluminium foil heat gently	pungent gas produced. Turns red litmus blue.
Sodium Carbonate	Flame test for sodium - yellow			
Iron(III) Chloride	$\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightarrow \text{Fe}(\text{OH})_3(\text{s})$	Same as with NaOH.	add nitric acid followed by silver nitrate	$\text{Ag}^{+}(\text{aq}) + \text{Cl}^{-}(\text{aq}) \rightarrow \text{AgCl}(\text{s})$
Sodium Bromide	Flame test for Sodium - yellow	_____		
Potassium Iodide	Flame test for Potassium - lilac	_____		
Aluminium Nitrate	Describe what happens	Describe what happens	Same as Calcium nitrate above	write equation. NO_3^{-} ions reduced to ammonia by the aluminium in alkaline conditions

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Page 1 of 2