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Page 1 of 1

Chapter 10 Electron Configuration and Periodicity

1. What is the definition of ionization energy?
The energy required to remove the outermost electron from a gaseous atom or ion.

2. What is the definition of electron affinity?
The energy change when one mole of gaseous atoms gains one mole of electrons to form one mole of gaseous anions.

3. What is the definition of electronegativity?
The ability of an atom in a molecule to attract the shared pair of electrons to itself.

4. What is the definition of covalent radius?
Half the distance between the nuclei of two covalently bonded atoms of the same element in a molecule.

5. What is the definition of ionic radius?
Half the distance between the nuclei of two ions in an ionic crystal.

6. What is the definition of atomic radius?
The radius of an atom.

7. What is the definition of crystal radius?
The radius of an ion in a crystal.

8. What is the definition of van der Waals radius?
Half the distance between the nuclei of two non-bonded atoms in a crystal.

9. What is the definition of effective nuclear charge?
The net positive charge experienced by an electron in a multi-electron atom.

10. What is the definition of shielding?
The reduction in the attractive force between the nucleus and the outermost electrons by the presence of other electrons.

11. What is the definition of penetration?
The ability of an electron to get close to the nucleus.

12. What is the definition of penetration energy?
The energy required to remove the most penetrating electron from an atom.

13. What is the definition of penetration effect?
The effect of the penetration of an electron into the nucleus.

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