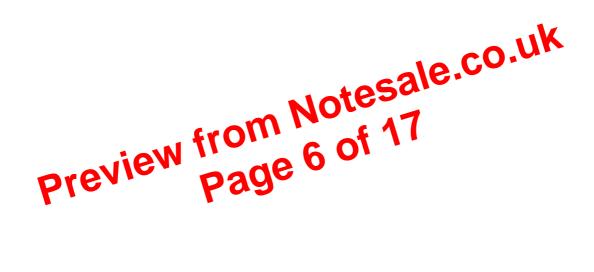
## **Review Questions**

- 1. Calculate the formula mass for each of the following.
  - a. K<sub>2</sub>SO<sub>4</sub>
  - b. CuO
  - c.  $Mg_3(AsO_4)_2$
  - d.  $Ca_3(PO_4)_2$
  - e. Fe<sub>2</sub>O<sub>3</sub>
  - f. Al(OH)<sub>3</sub>
  - g.  $(NH_4)_2S$
  - h.  $C_{12}H_{22}O_{11}$
- 2. On average, how many times heavier are bromine atoms than neon atoms?
- 3. An unknown element, M, combines with oxygen to form a compound with a formula of  $MO_2$ . If 25.0 grams of the unknown element combines with 4.50 grams of oxygen, what is the atomic mass of M?



## **Lesson Summary**

- The percent composition of a compound is the percent of the total mass contributed by each element in the compound.
- Percent composition can be determined either from the masses of each element in the compound or from the formula of the compound.

## **Further Reading / Supplemental Links**

This website has solved example problems for a number of topics covered in this lesson, including the calculation of percent composition by mass.

• http://www.sciencejoywagon.com/chemzone/05chemical-reactions/

This website has lessons, worksheets, and quizzes on various high school chemistry topics. Lesson 5-8 is on percent composition.

• http://www.fordhamprep.org/gcurran/sho/sho/lessons/lesson58.htm

Review Questions eV

Determine the percent composition of the 15 howing act

- 2.  $Ca(C_2H_3O_2)_2$
- 3. FeF<sub>3</sub>
- 4. CrCl<sub>3</sub>
- 5. (NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub>