

Molarity

$$M = \frac{\text{mol}}{L}$$

exy

$$HCL = \frac{2 \text{ mol}}{L} = \frac{73 \text{ g}}{L}$$

$$\text{g C} \times \frac{1 \text{ mol}}{12 \text{ g C}} = \text{mol C}$$

$$H_2O = 1 \text{ g/cm}^3 \text{ or } 1 \text{ g/mL}$$

$$73 \text{ g} = ? \%$$

$$\frac{73}{1073} = \frac{x}{100} = 6.8 \%$$

$$6.8/100$$

PPH

$$68/1,000$$

PPt

$$68,000/1,000,000$$

PPM

$$68,000/1,000,000$$

PPB

$$1 \text{ mol H}_2\text{O} = 18 \text{ g}$$

$$3 \text{ mol C} = 36 \text{ g}$$

$$2 \text{ mol C} = 24 \text{ g}$$

$$1 \text{ mol Li} = 7 \text{ g}$$

$$2 \text{ mol Li} = 1.2 \times 10^{24} \text{ g}$$

$$18 \text{ g H}_2\text{O} = 1 \text{ mol}$$

$$3.6 \times 10^{24} \text{ molecules C}_6\text{H}_{12}\text{O}_6 = 4 \text{ mol}$$

$$C = 12 \text{ g/mol}$$

$$6 \text{ mol C} \times \frac{12 \text{ g C}}{1 \text{ mol}} = 72 \text{ g C}$$