

PH < 7 is acidity

PH = 7 neutral

PH > 7 is basic

PH = $-\log[H^+]$

PH = $-\log$

Same as concentration of hydroxide ions

PH = $14 - 7 = 7$

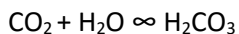
Acid Deposition

Wet deposition – acid rain (including fog, snow and dew)

Dry deposition – acidic gases and particles

PH of rainwater = 5.65

Slightly acidic due to dissolved carbon dioxide



Main gases causing acid deposition:

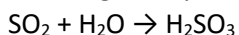
Oxides of Sulphur

Oxides of nitrogen

SO₂ : volcanoes

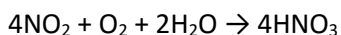
Combustion of Sulphur containing fossil fuels

Smelting of sulphide ores



NO_x

Occurs in electrical storms and bacterial action



Effects of acid rain:

1. Vegetation

Many trees have been affected by the rain

2. Lakes and rivers

Below pH 4 lakes are dead

Presence of nitrates present in acid rain can also lead to eutrophication

3. Buildings:

