Catalyst- increase the rate of reaction by lowering the Activation Energy because more successful collisions because more particles will have collided with the Activation Energy.

Mole calculations 6x10 power23 n сν m n mr v dm3 n 24 dm3 C fl cl 0.24 0.38 1.42 12 19 35 0.02 0.02 0.04 2 1 1 C2F2Cl4 Mr= 204 esale.co.uk dm-cm= Times by 1000 cm-dm= Divide by 1000 **Electrolysis** Current- The flow of electrons or ions. They will only conduct electricity when they are n Electrolyte- Substances which on ucts electricity when moliten or dissolved in water, ionic 20 compounds are electrol Metals- Cations surrounded by delocalised electrons.

Electrolysis- Involves the formation of a new substance when ionic compounds conduct electricity.

Cathode (-) Positive ions go to cathode (Metals + Hydrogen)- Collect them. Anode (+) Negative ions go to Anode (Non- Metals)- Dump them.

Molten Ionic Substances PbBr2

Cation(+) Pb2 + $2e \rightarrow Pb$ (Cathode) Anode(-) 2Br- $\rightarrow Br2 + 2e$ (Anode)