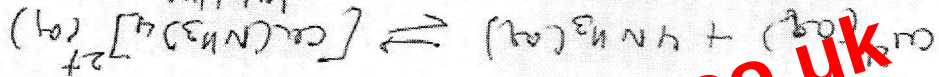


H_3PO_2 has one ionizable P-OH bond while
 H_2PO_3 has two ionizable P-OH bonds & H_2PO_4 has
 3 ionizable P-OH bonds as H_3PO_3 is monobasic, H_2PO_3
 is dibasic and H_2PO_4 is tribasic.

In H_3PO_2 , two H atoms are bonded directly
 to P atom which impart reducing character to
 the acid. While H_2PO_3 has one H atom directly
 bonded to P atom making it comparatively weaker reducing agent.
 Both are sp^3 hybridized, in PH_3 all the 4 orbitals
 are bonded whereas in PH_3 there is a lone pair of
 electrons on P, which is responsible for the lone pair - bond
 pair repulsion in PH_3 reducing the bond angle to less
 than $109^\circ 28'$.

Other types of greenhouse gases

High pressure & low temperature cause of a catalyst such
 as iron oxide & small amounts of K_2O and Al_2O_3 .



Colorance of nitrogen in N_2O_5 is $\frac{4}{5}$ deep blue

Brown ring test $\Rightarrow$ Aqueous solution of nitrate + freshly prepared
 ferrous sulphate solution & then carefully conc. H_2SO_4 is
 added along the sides of the test. A brown ring is
 formed at the junction of 2 liquids.

Reason: $9t$ is due to formation of $[Fe(H_2O)_5(NO)]^{2+}$

White phosphorus	Red phosphorus
1. White phosphorus is translucent white waxy add.	1. $9t$ persists in grey shales.
2. It is poisonous	2. It is non poisonous
3. It is insoluble in H_2O but soluble in CS_2 .	3. $9t$ is insoluble both in H_2O & CS_2 .