

Hess's Law

27 December 2012
10:46

Formulas & Theorems Covered Today:

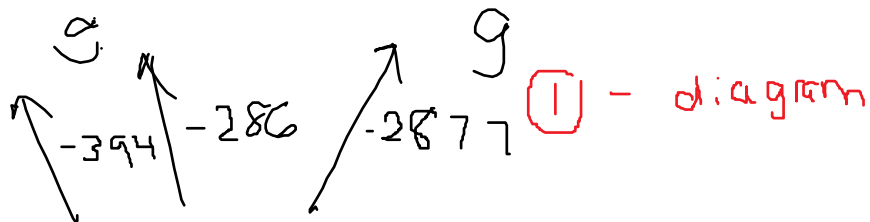
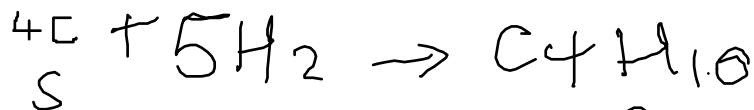
★ $\Delta H = \sum \text{products} - \sum \text{reactants}$



EXAMPLE:

Find the ΔH_f of butane given the following data:

ΔH_c : $C_4H_{10}(g) = -104$, $C(s) = -394$, $H_2(g) = -286 \text{ kJ mol}^{-1}$



$$-104 - (394 \times 4 + (286 \times 5))$$

$$= -125 \text{ kJ mol}^{-1}$$

Preview from Notesale.co.uk
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① units
① answer
4/4.